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Chapter 6  
Chemical  
Bonding

# **Chapter 6 Review Chemical Bonding**

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Atomic Hook-Ups -  
Types of Chemical  
Bonds: Crash  
Course Chemistry  
#22 Bonding*

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*Models and Lewis Structures: Crash Course Chemistry*

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### **Chemistry - 6 - Bonding and Structure 6.1**

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~~Chemical Bonding Section 1 \u0026amp; 2~~

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Chemical Bonding

Covalent Bonds

and Ionic Bonds

~~Chemical Bonding~~

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*Covalent Bonding |*

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8 Basic Concepts of  
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**VSEPR Theory:**

**Introduction**

*Chemical Bonding -*

*Ionic vs. Covalent*

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Chemistry -

Covalent Bonding

#14

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6.2 Covalent

Bonding and

Molecular

Compounds Ionic

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Bonds Made Easy

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Lewis Dot

Structures **Lewis**

**Structures,**

**Introduction,**

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**Formal Charge,  
Molecular  
Geometry,  
Resonance, Polar  
or Nonpolar**

*Orbitals: Crash*

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**#25 Ionic and**

**Covalent Bonds,  
Hydrogen Bonds,  
van der Waals - 4**

**types of**

**Chemical Bonds  
in Biology**

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*Chemical Bonding*

Types Of Chemical

Bonds - What Are

Chemical Bonds -

Covalent Bonds

And Ionic Bonds -

What Are Ions

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Chemical Bond?

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CHAPTER 6 REVIEW

Chemical Bonding

SECTION 1 SHORT

ANSWER Answer

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the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent

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bond consists of (a)  
a shared electron.

### *6 Chemical Bonding*

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Chemical bond  
between atoms  
results from the  
attraction between  
the valence  
electrons and \_\_\_\_\_  
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nuclei. A covalent  
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Bonding SECTION 1

SHORT ANSWER

Answer the following questions in the space provided. 1. \_\_\_\_\_ A chemical bond between atoms results from the attraction between the valence electrons and \_\_\_\_\_ of different atoms. (a) nuclei

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(c) isotopes (b)  
inner electrons (d)  
Lewis structures 2.

## Bonding

*CHAPTER 6 REVIEW*

*Chemical Bonding*

CHAPTER 6 REVIEW

Chemical Bonding

SECTION 5 SHORT

ANSWER Answer

the following

questions in the

space provided. 1.

Identify the major

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assumption of the VSEPR theory, which is used to predict the shape of atoms. Pairs of valence electrons repel one another.

2. In water, two hydrogen atoms are bonded to one oxygen atom. Why isn't water a linear molecule?

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Chemical bonding that results from the electrical attraction between positive ions and negative ions is called ionic bonding. In a purely ionic bond, the metal atom gives up its

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electron or  
electrons to the  
nonmetal atom. In  
covalent bonding, a  
bond forms from  
the sharing of  
electron pairs  
between two  
atoms. In a purely  
covalent bond, the

*CHAPTER 6*

*Chemical Bonding*

Section 6.2 –

*Page 24/37*



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Covalent Bonding A

covalent bond is a chemical bond in which two atoms share a pair of valence electrons.

When two atoms share one pair of electrons, the bond is called a single bond. Covalent vs Ionic Bond

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*Page 25/37*

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### *Chemical Bonds*

an attraction between atoms that allows the formation of chemical substances that contain two or more atoms. The bond is caused by the electrostatic force of attraction between opposite charges, either

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between electrons and nuclei, or as the result of a dipole attraction.

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bonding that

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Chemical Bonding

SECTION 6-2

SHORT ANSWER

Answer the

following questions

in the space

provided. 1. Use

the concept of

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potential energy to describe how a covalent bond forms between two atoms. 2. Name two elements that form compounds that are exceptions to the octet rule. 3.

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Molecules. How many electrons should Helium have around its Lewis dot model?

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ionic bonds,  
covalent bonds,

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and metallic bonds.

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Lewis dot  
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structures, and  $\hat{\epsilon}$ !

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