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reduction (c) cathode, oxidation (d)  
cathode, reduction

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If it displaces  $\text{Au}^+ (\text{aq})$  from solution,  
then it has a reduction potential  
smaller than  $E^\circ_{\text{Au}^+/\text{Au}}$

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1.68V. But if it does not displace  $\text{Fe}^{3+}(\text{aq})$  from solution, then its reduction potential is larger than  $E^\circ(\text{Fe}^{3+}(\text{aq})/\text{Fe}^{2+}(\text{aq})) = 0.769\text{V}$ . Therefore,  $0\text{V} < E^\circ < 0.17\text{V}$ .

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Problems. 1). Given the  $E^\circ$  for the

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following half-reactions:  $\text{Cu}^+ + e^- \rightarrow \text{Cu}^\circ$  ?

$\text{Cu}^{2+} + 2e^- \rightarrow \text{Cu}^\circ$  ?

$\text{Cu}^\circ$  ?

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Answer. Using rule 5 and 7.

$\text{MgF}_2$  total charge=0 Total

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$$\text{Charge} = (+2) + (-1 \times 2) = 0$$

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An electrochemical cell affords us a high degree of control and measurement of the cell reaction. If the external circuit is broken, the

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reaction stops. If we place a variable resistance in the circuit, we can control the rate of the cell reaction by simply turning a knob. By connecting a battery or other source of

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Answer/Explanation. Answer: b.

Explanation: (b) It is because  $E^\circ = 0.77$  V, it means  $\text{Fe}^{3+}$  can gain electron early to form  $\text{Fe}^{2+}$ . 10. ? of M/32 solution of weak acid is  $8 \text{ S cm}^2 \text{ mol}^{-1}$  and limiting molar conductivity is  $400 \text{ S cm}^2 \text{ mol}^{-1}$ .  $K_a$  for acid is. (a)  $1.25 \times 10^{-6}$ . (b)  $6.25 \times 10^{-4}$ .

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The electrochemical cell which will involve the given cell reaction is depicted as  $Z_n / Z_n^{2+}(aq) || Ag^+(aq) / Ag$  (i) In this cell, the  $Z_n / Z_n^{2+}$  will act as negative

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terminal. and  $A g / A g^{+}$  will act as positive terminal (ii) The conventional current will flow from silver to zinc electrode in the external circuit

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10. Molten NaCl conducts electricity due to the presence of. A. Free Electrons. B. Free molecules. C. Free ions. D. Atoms of Na and Cl. View Answer. Answer: Option C. Click On Below Page Numbers To Go To Next/Previous Page of Electrochemistry MCQs.



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An electrochemical cell is shown below  $\text{Pt, H}_2 (1 \text{ atm}) | \text{HCl} (0.1 \text{ M}) | \text{CH}_3\text{COOH} (0.1 \text{ M}) | \text{H}_2 (1 \text{ atm}), \text{Pt}$  The EMF of the cell will not be zero, because (a) EMF depends on molarities of acids used (b) pH of 0.1

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M HCl and 0.1 M CH<sub>3</sub>COOH is not same (c) the temperature is constant (d) acids used in two compartments are different

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identifiable chemical change.

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The branch of chemistry that deals with the study of redox reactions and how they can be applied to generate electricity (in electrochemical cells) and to carry out non-spontaneous reactions using electricity (electrolysis).

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(s)/Zn<sup>2+</sup> (aq) 1M || Cu<sup>2+</sup> (aq) 1M/Cu

(s) is representation of reaction in. A. Daniel cell. B. Downs cell. C. Voltaic cell. D. Nelsons cell. View Answer.

Answer: Option C.

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